metal-organic compounds

catena-Poly[[bis(3,5-dicarboxybenzo-ato)cobalt(II)]-*μ*-4,4′-bipyridine]

Chun-Sheng Ling[a] and Qiu-Xia Han[b]*

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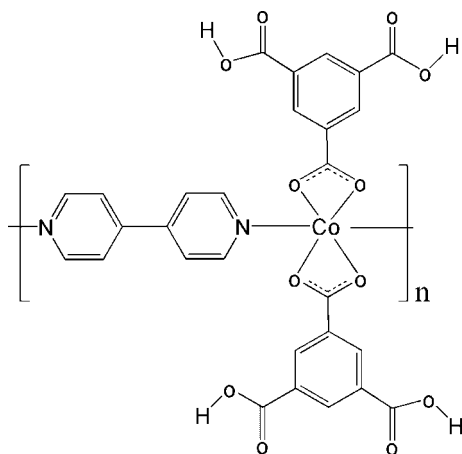Correspondence e-mail: hdhqx@henu.edu.cn

In the title compound, $[Co(C_9H_5O_6)_2(C_{10}H_8N_2)]_n$, the asymmetric unit consists of one Co^{2+} ion with site symmetry 2, one mono-deprotonated 1,3,5-benzenetricarboxylic acid anion and one-half of a 4,4′-bipyridine (4,4′-bipy) molecule, in which two N and two C atoms have site symmetry 2. In the crystal structure, the Co^{2+} centre is coordinated by four O atoms from two bidentate carboxylate groups of two anions and two N atoms of two 4,4′-bipy molecules, resulting in infinite chains propagating in [010]. The cobalt coordination is distorted *trans*-CoO_4N_2 octahedral and interchain O—H···O hydrogen bonds complete the structure.

Related literature

For background, see: Feller *et al.* (2007); Brown *et al.* (2008).

Experimental

Crystal data

$[Co(C_9H_5O_6)_2(C_{10}H_8N_2)]$	V = 2617.9 (3) Å^3
M_r = 633.37	Z = 4
Monoclinic, $C2/c$	Mo $K\alpha$ radiation
a = 10.6682 (7) Å	μ = 0.73 mm^{-1}
b = 11.0490 (7) Å	T = 293 (2) K
c = 22.6563 (14) Å	$0.18 \times 0.15 \times 0.13$ mm
β = 101.401 (1)°	

Data collection

Bruker SMART CCD diffractometer	7123 measured reflections
Absorption correction: multi-scan (*SADABS*; Sheldrick, 2001)	2579 independent reflections
T_{min} = 0.880, T_{max} = 0.911	2051 reflections with $I > 2\sigma(I)$
	R_{int} = 0.031

Refinement

$R[F^2 > 2\sigma(F^2)]$ = 0.038	7 restraints
$wR(F^2)$ = 0.098	H-atom parameters constrained
S = 1.02	$\Delta\rho_{max}$ = 0.30 e Å^{-3}
2579 reflections	$\Delta\rho_{min}$ = −0.73 e Å^{-3}
197 parameters	

Table 1
Selected bond lengths (Å).

Co1—N2^i	1.982 (2)	Co1—O1	2.0221 (14)
Co1—N1	1.992 (2)	Co1—O2	2.4354 (13)

Symmetry code: (i) $x, y + 1, z$.

Table 2
Hydrogen-bond geometry (Å, °).

D—H···A	D—H	H···A	D···A	D—H···A
O3—H3···O2^{ii}	0.81	1.88	2.648 (2)	157
O5—H5···O6^{iii}	0.78	1.88	2.651 (2)	170

Symmetry codes: (ii) $-x + \frac{1}{2}, -y + \frac{3}{2}, -z$; (iii) $-x + 2, -y + 1, -z$.

Data collection: *SMART* (Bruker, 2001); cell refinement: *SAINT-Plus* (Bruker, 2001); data reduction: *SAINT-Plus*); program(s) used to solve structure: *SHELXS97* (Sheldrick, 2008); program(s) used to refine structure: *SHELXL97* (Sheldrick, 2008); molecular graphics: *PLATON* (Spek, 2003); software used to prepare material for publication: *PLATON*.

References

Brown, K. A., Martin, D. P. & LaDuca, R. L. (2008). *CrystEngComm*, **10**, 1305–1308.
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supplementary materials

catena-Poly[[bis(3,5-dicarboxybenzoato)cobalt(II)]-*μ*-4,4'-bipyridine]

C.-S. Ling and Q.-X. Han

Comment

Recently, many efforts in coordination chemistry and crystal engineering have been devoted to the construction of metal-organic coordination polymers (MOCPs) employing both coordination bonds and/or hydrogen bonds, due to their appropriate strength and directionality (Feller *et al.* 2007). Dual-ligand or multidentate organic ligands are usually engaged in the construction of MOCPs, among which carboxylates and N,*N*-bidentate ligands are all the simplest connectors potentially able to bridge metal ions (Brown *et al.* 2008). Herein, we report the title compound (I) containing organic dual-ligands (Fig. 1).

The structure of (I) presents a one-dimensional infinite chain (Fig.2), in which the Co^{2+} centre (site symmetry 2) is coordinated by four O atoms from two bidentate carboxylate groups of two 1,3,5-benzenetricarboxylic acid anions, two N atoms of two 4,4'-bipyridine molecules. The Co^{2+} caion resides in a distorted octahedral configuration. In the equatorial plane, it is chelted by four carboxylate oxygen atoms (O1, O2 and their symmetry equivalents) from two 1,3,5-benzenetricarboxylic acid anions (Table 1), in which the Co—O distances are very different.

In addition, these one-dimensional chains are linked together by O—H···O hydrogen bonds between carboxylate groups generating a three-dimensional framework (Fig. 3 and Table 2).

Experimental

Solid $Co(CH_3COO)_2.4H_2O$ (1 mmol, 0.245 g) was added to an aqueous solution (25 ml) of 1,3,5-benzenetricarboxylic acid (2 mmol, 0.420 g) and 4,4'-bipyridine (1 mmol, 0.156 g). The mixture was refluxed for two hours at 373 K. The solution was filtered, and the filtrate was kept at room temperature. After ten days, purple blocks of (I) were obtained.

Refinement

The O-bound H atoms were located in difference Fourier maps and refined as riding in their as-found relative positions with $U_{iso}(H) = 1.5U_{eq}(O)$. The C-bound H atoms were geometrically placed (C—H = 0.93Å) and refined as riding, $U_{iso}(H) = 1.2U_{eq}(C)$.
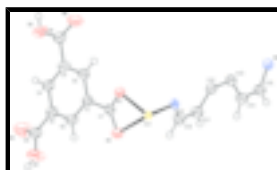
Figures

Fig. 1. Asymmetric unit of (I), showing displacement ellipsoids at the 50% probability level for the non-hydrogen atoms.
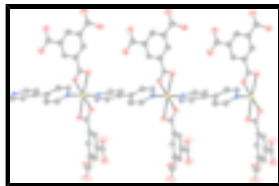
supplementary materials

Fig. 2. One-dimensional chain structure of (I). H atoms are omitted for clarity. Displacement ellipsoids are drawn at the 50% probability level.
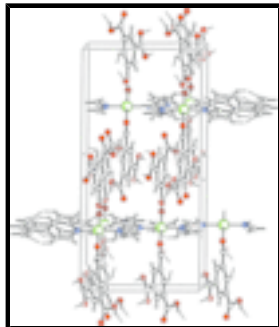

Fig. 3. Three-dimensional structure of (I) arising by means of hydrogen bonds. Displacement ellipsoids are drawn at the 50% probability level.

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Crystal data

$[Co(C_9H_5O_6)_2(C_{10}H_8N_2)]$	$F_{000} = 1292$
$M_r = 633.37$	$D_x = 1.607$ Mg m^{-3}
Monoclinic, $C2/c$	Mo $K\alpha$ radiation $\lambda = 0.71073$ Å
Hall symbol: -C 2yc	Cell parameters from 1750 reflections
$a = 10.6682$ (7) Å	$\theta = 2.7–25.9º$
$b = 11.0490$ (7) Å	$\mu = 0.73$ mm^{-1}
$c = 22.6563$ (14) Å	$T = 293$ (2) K
$\beta = 101.401$ (1)º	Block, purple
$V = 2617.9$ (3) Å^3	$0.18 \times 0.15 \times 0.13$ mm
$Z = 4$	

Data collection

Bruker SMART CCD diffractometer	2579 independent reflections
Radiation source: fine-focus sealed tube	2051 reflections with $I > 2\sigma(I)$
Monochromator: graphite	$R_{int} = 0.031$
$T = 293(2)$ K	$\theta_{max} = 26.0º$
ω scans	$\theta_{min} = 2.7º$
Absorption correction: multi-scan (SADABS; Sheldrick, 2001)	$h = -13 \rightarrow 13$
$T_{min} = 0.880$, $T_{max} = 0.911$	$k = -13 \rightarrow 13$
7123 measured reflections	$l = -27 \rightarrow 15$

Refinement

Refinement on F^2	Secondary atom site location: difference Fourier map

Least-squares matrix: full

$R[F^2 > 2\sigma(F^2)] = 0.038$

$wR(F^2) = 0.098$

$S = 1.02$

2579 reflections

197 parameters

7 restraints

Primary atom site location: structure-invariant direct methods

Hydrogen site location: difmap and geom

H-atom parameters constrained

$w = 1/[\sigma^2(F_o^2) + (0.0508P)^2 + 1.9201P]$
where $P = (F_o^2 + 2F_c^2)/3$

$(\Delta/\sigma)_{max} = 0.001$

$\Delta\rho_{max} = 0.31$ e Å^{-3}

$\Delta\rho_{min} = -0.73$ e Å^{-3}

Extinction correction: none

Special details

Geometry. All e.s.d.'s (except the e.s.d. in the dihedral angle between two l.s. planes) are estimated using the full covariance matrix. The cell e.s.d.'s are taken into account individually in the estimation of e.s.d.'s in distances, angles and torsion angles; correlations between e.s.d.'s in cell parameters are only used when they are defined by crystal symmetry. An approximate (isotropic) treatment of cell e.s.d.'s is used for estimating e.s.d.'s involving l.s. planes.

Refinement. Refinement of F^2 against ALL reflections. The weighted R-factor wR and goodness of fit S are based on F^2, conventional R-factors R are based on F, with F set to zero for negative F^2. The threshold expression of $F^2 > \sigma(F^2)$ is used only for calculating R-factors(gt) *etc*. and is not relevant to the choice of reflections for refinement. R-factors based on F^2 are statistically about twice as large as those based on F, and R- factors based on ALL data will be even larger.

Fractional atomic coordinates and isotropic or equivalent isotropic displacement parameters (Å^2)

	x	y	z	U_{iso}*/U_{eq}
Co1	0.5000	0.64801 (3)	0.2500	0.01788 (9)
O1	0.60733 (12)	0.64666 (13)	0.18575 (6)	0.0280 (3)
O2	0.40351 (13)	0.65638 (14)	0.14327 (6)	0.0347 (4)
O3	0.30786 (17)	0.7752 (2)	−0.06851 (8)	0.0661 (6)
H3	0.2523	0.7863	−0.0982	0.099*
O4	0.39825 (18)	0.6736 (2)	−0.13349 (8)	0.0692 (6)
O5	0.85792 (15)	0.55679 (17)	−0.04477 (7)	0.0509 (5)
H5	0.9287	0.5390	−0.0439	0.076*
O6	0.91135 (16)	0.5199 (2)	0.05406 (8)	0.0591 (6)
N1	0.5000	0.4678 (2)	0.2500	0.0232 (4)
N2	0.5000	−0.1726 (2)	0.2500	0.0274 (6)
C1	0.55508 (19)	0.64578 (18)	0.07857 (9)	0.0281 (5)
C2	0.67509 (19)	0.60572 (19)	0.07160 (9)	0.0298 (5)
H2	0.7360	0.5843	0.1053	0.036*
C3	0.70461 (19)	0.5975 (2)	0.01464 (10)	0.0306 (5)
C4	0.6137 (2)	0.6293 (2)	−0.03605 (10)	0.0338 (5)
H4	0.6326	0.6222	−0.0742	0.041*
C5	0.4946 (2)	0.6717 (2)	−0.02928 (10)	0.0329 (5)
C6	0.4664 (2)	0.6797 (2)	0.02785 (10)	0.0327 (5)
H6	0.3867	0.7084	0.0322	0.039*
C7	0.51892 (19)	0.65013 (18)	0.13895 (9)	0.0265 (5)
C8	0.8332 (2)	0.5551 (2)	0.00808 (10)	0.0362 (6)

C9	0.3971 (2)	0.7059 (2)	−0.08302 (10)	0.0392 (6)
C10	0.4034 (2)	0.40519 (19)	0.21766 (10)	0.0321 (5)
H10	0.3356	0.4477	0.1947	0.039*
C11	0.3993 (2)	0.28136 (19)	0.21669 (10)	0.0330 (5)
H11	0.3295	0.2414	0.1938	0.040*
C12	0.5000	0.2156 (3)	0.2500	0.0284 (7)
C13	0.5000	0.0812 (3)	0.2500	0.0269 (7)
C14	0.38822 (19)	0.01470 (19)	0.23157 (10)	0.0333 (5)
H14	0.3108	0.0547	0.2192	0.040*
C15	0.3916 (2)	−0.1093 (2)	0.23148 (10)	0.0328 (5)
H15	0.3158	−0.1514	0.2181	0.039*

Atomic displacement parameters (Å^2)

	U^{11}	U^{22}	U^{33}	U^{12}	U^{13}	U^{23}
Co1	0.02150 (17)	0.01328 (16)	0.01973 (17)	0.000	0.00619 (13)	0.000
O1	0.0246 (6)	0.0345 (8)	0.0250 (7)	0.0022 (6)	0.0052 (5)	0.0008 (6)
O2	0.0252 (7)	0.0480 (9)	0.0314 (8)	0.0040 (7)	0.0071 (6)	0.0026 (7)
O3	0.0446 (10)	0.1163 (17)	0.0354 (10)	0.0378 (11)	0.0033 (8)	0.0035 (10)
O4	0.0557 (11)	0.1214 (18)	0.0293 (10)	0.0270 (12)	0.0052 (8)	−0.0022 (10)
O5	0.0367 (8)	0.0734 (12)	0.0478 (10)	0.0152 (8)	0.0210 (7)	0.0025 (9)
O6	0.0376 (9)	0.0938 (15)	0.0466 (10)	0.0261 (9)	0.0100 (8)	0.0094 (10)
N1	0.0263 (7)	0.0191 (7)	0.0246 (7)	0.000	0.0062 (6)	0.000
N2	0.0268 (12)	0.0246 (13)	0.0313 (13)	0.000	0.0072 (10)	0.000
C1	0.0268 (10)	0.0300 (10)	0.0276 (10)	0.0007 (9)	0.0057 (8)	−0.0005 (9)
C2	0.0250 (10)	0.0327 (11)	0.0308 (11)	0.0034 (9)	0.0030 (9)	0.0013 (9)
C3	0.0254 (10)	0.0333 (11)	0.0335 (11)	0.0015 (9)	0.0070 (9)	−0.0025 (9)
C4	0.0294 (10)	0.0441 (13)	0.0296 (11)	0.0011 (10)	0.0097 (9)	−0.0015 (10)
C5	0.0289 (11)	0.0410 (13)	0.0288 (11)	0.0011 (9)	0.0054 (9)	0.0003 (9)
C6	0.0256 (10)	0.0413 (12)	0.0320 (12)	0.0031 (9)	0.0071 (9)	−0.0010 (10)
C7	0.0259 (10)	0.0244 (10)	0.0291 (10)	0.0014 (8)	0.0050 (8)	0.0002 (9)
C8	0.0297 (11)	0.0435 (13)	0.0369 (13)	0.0022 (10)	0.0100 (10)	−0.0008 (10)
C9	0.0292 (11)	0.0614 (15)	0.0284 (12)	0.0044 (11)	0.0092 (9)	0.0027 (11)
C10	0.0317 (11)	0.0265 (11)	0.0370 (12)	0.0027 (9)	0.0042 (9)	0.0031 (9)
C11	0.0310 (11)	0.0271 (11)	0.0391 (12)	−0.0005 (9)	0.0030 (9)	−0.0002 (9)
C12	0.0267 (14)	0.0246 (15)	0.0354 (16)	0.000	0.0096 (12)	0.000
C13	0.0270 (14)	0.0238 (15)	0.0300 (16)	0.000	0.0057 (12)	0.000
C14	0.0246 (10)	0.0265 (11)	0.0464 (13)	0.0007 (9)	0.0015 (10)	0.0022 (10)
C15	0.0241 (10)	0.0276 (11)	0.0456 (13)	−0.0015 (9)	0.0044 (10)	0.0006 (10)

Geometric parameters (Å, °)

Co1—N2^i	1.982 (2)	C1—C7	1.494 (3)
Co1—N1	1.992 (2)	C2—C3	1.390 (3)
Co1—O1^{ii}	2.0221 (14)	C2—H2	0.9300
Co1—O1	2.0221 (14)	C3—C4	1.393 (3)
Co1—O2^{ii}	2.4354 (13)	C3—C8	1.485 (3)
Co1—O2	2.4354 (13)	C4—C5	1.391 (3)

O1—C7	1.273 (2)	C4—H4	0.9300
O2—C7	1.256 (2)	C5—C6	1.389 (3)
O3—C9	1.314 (3)	C5—C9	1.485 (3)
O3—H3	0.8127	C6—H6	0.9300
O4—C9	1.200 (3)	C10—C11	1.369 (3)
O5—C8	1.276 (3)	C10—H10	0.9300
O5—H5	0.7761	C11—C12	1.390 (3)
O6—C8	1.260 (3)	C11—H11	0.9300
N1—C10	1.333 (2)	C12—C11^{ii}	1.390 (3)
N1—C10^{ii}	1.333 (2)	C12—C13	1.485 (4)
N2—C15^{ii}	1.346 (2)	C13—C14	1.392 (2)
N2—C15	1.346 (2)	C13—C14^{ii}	1.392 (2)
N2—Co1^{iii}	1.982 (2)	C14—C15	1.371 (3)
C1—C6	1.388 (3)	C14—H14	0.9300
C1—C2	1.393 (3)	C15—H15	0.9300
N2^{i}—Co1—N1	180.0	C1—C6—C5	121.1 (2)
N2^{i}—Co1—O1^{ii}	90.42 (4)	C1—C6—H6	119.5
N1—Co1—O1^{ii}	89.58 (4)	C5—C6—H6	119.5
N2^{i}—Co1—O1	90.42 (4)	O2—C7—O1	120.87 (19)
N1—Co1—O1	89.58 (4)	O2—C7—C1	120.51 (18)
O1^{ii}—Co1—O1	179.16 (8)	O1—C7—C1	118.62 (17)
C7—O1—Co1	99.65 (12)	O6—C8—O5	123.7 (2)
C9—O3—H3	109.1	O6—C8—C3	119.2 (2)
C8—O5—H5	110.6	O5—C8—C3	117.11 (19)
C10—N1—C10^{ii}	117.5 (2)	O4—C9—O3	123.8 (2)
C10—N1—Co1	121.23 (12)	O4—C9—C5	124.6 (2)
C10^{ii}—N1—Co1	121.23 (12)	O3—C9—C5	111.60 (19)
C15^{ii}—N2—C15	117.4 (2)	N1—C10—C11	123.1 (2)
C15^{ii}—N2—Co1^{iii}	121.30 (12)	N1—C10—H10	118.5
C15—N2—Co1^{iii}	121.30 (12)	C11—C10—H10	118.5
C6—C1—C2	118.9 (2)	C10—C11—C12	119.7 (2)
C6—C1—C7	119.49 (18)	C10—C11—H11	120.2
C2—C1—C7	121.58 (18)	C12—C11—H11	120.2
C3—C2—C1	120.52 (19)	C11^{ii}—C12—C11	116.9 (3)
C3—C2—H2	119.7	C11^{ii}—C12—C13	121.54 (13)
C1—C2—H2	119.7	C11—C12—C13	121.54 (13)
C2—C3—C4	120.00 (19)	C14—C13—C14^{ii}	116.3 (3)
C2—C3—C8	119.79 (18)	C14—C13—C12	121.83 (13)
C4—C3—C8	120.2 (2)	C14^{ii}—C13—C12	121.83 (13)
C5—C4—C3	119.7 (2)	C15—C14—C13	120.4 (2)
C5—C4—H4	120.1	C15—C14—H14	119.8
C3—C4—H4	120.1	C13—C14—H14	119.8
C6—C5—C4	119.72 (19)	N2—C15—C14	122.7 (2)
C6—C5—C9	120.1 (2)	N2—C15—H15	118.7
C4—C5—C9	120.2 (2)	C14—C15—H15	118.7

N2^{i}—Co1—O1—C7	−88.18 (11)	C6—C1—C7—O1	−163.76 (19)
N1—Co1—O1—C7	91.82 (11)	C2—C1—C7—O1	17.7 (3)
O1^{ii}—Co1—O1—C7	91.82 (12)	C2—C3—C8—O6	4.3 (3)
N2^{i}—Co1—N1—C10	−141 (22)	C4—C3—C8—O6	−175.8 (2)
O1^{ii}—Co1—N1—C10	83.93 (12)	C2—C3—C8—O5	−175.6 (2)
O1—Co1—N1—C10	−96.07 (12)	C4—C3—C8—O5	4.3 (3)
N2^{i}—Co1—N1—C10^{ii}	39 (23)	C6—C5—C9—O4	−159.5 (3)
O1^{ii}—Co1—N1—C10^{ii}	−96.07 (12)	C4—C5—C9—O4	19.2 (4)
O1—Co1—N1—C10^{ii}	83.93 (12)	C6—C5—C9—O3	19.8 (3)
C6—C1—C2—C3	−1.3 (3)	C4—C5—C9—O3	−161.5 (2)
C7—C1—C2—C3	177.20 (19)	C10^{ii}—N1—C10—C11	0.39 (16)
C1—C2—C3—C4	−0.1 (3)	Co1—N1—C10—C11	−179.61 (16)
C1—C2—C3—C8	179.8 (2)	N1—C10—C11—C12	−0.8 (3)
C2—C3—C4—C5	1.4 (3)	C10—C11—C12—C11^{ii}	0.37 (15)
C8—C3—C4—C5	−178.4 (2)	C10—C11—C12—C13	−179.63 (15)
C3—C4—C5—C6	−1.3 (3)	C11^{ii}—C12—C13—C14	161.59 (15)
C3—C4—C5—C9	180.0 (2)	C11—C12—C13—C14	−18.41 (15)
C2—C1—C6—C5	1.5 (3)	C11^{ii}—C12—C13—C14^{ii}	−18.41 (15)
C7—C1—C6—C5	−177.08 (19)	C11—C12—C13—C14^{ii}	161.59 (15)
C4—C5—C6—C1	−0.2 (3)	C14^{ii}—C13—C14—C15	−0.70 (16)
C9—C5—C6—C1	178.6 (2)	C12—C13—C14—C15	179.30 (16)
Co1—O1—C7—O2	1.9 (2)	C15^{ii}—N2—C15—C14	−0.74 (16)
Co1—O1—C7—C1	−177.68 (15)	Co1^{iii}—N2—C15—C14	179.26 (16)
C6—C1—C7—O2	16.6 (3)	C13—C14—C15—N2	1.5 (3)
C2—C1—C7—O2	−161.9 (2)		

Symmetry codes: (i) x, $y+1$, z; (ii) $-x+1$, y, $-z+1/2$; (iii) x, $y-1$, z.

Hydrogen-bond geometry (Å, °)

D—H⋯A	D—H	H⋯A	D⋯A	D—H⋯A
O3—H3⋯O2^{iv}	0.81	1.88	2.648 (2)	157
O5—H5⋯O6^{v}	0.78	1.88	2.651 (2)	170

Symmetry codes: (iv) $-x+1/2$, $-y+3/2$, $-z$; (v) $-x+2$, $-y+1$, $-z$.

Fig. 1

Fig. 2

Fig. 3